

# NOTES # 6/Aqueous Chem B/Reactions/AP Chemistry

## 1. Observable Effects of Rxns (How do we know something happened?)

- Precipitates
- Gases
- Formation of electrolyte
- $\Delta$  oxidation#
- Breaking covalent bonds
- Change in appearance/color
- Explosions/light emission

## 2. Types of reactions

### Precipitate Forming RXNS:

General Equation:

-Atoms/Ions switch partners

-An insoluble product or precipitate MUST BE FORMED!

- Memorize the solubility chart on page 118 in Brown, LeMay, Bursten!!!

Soluble Ionic Compounds (aq)  
(ALWAYS soluble)

- Insoluble Ionic Compounds (s)  
(usually insoluble, unless with ALWAYS)

## **Writing Net Ionic Equations:**

- Determine Products
- Determine Solubility
- Will the rxn proceed?
- Write the Ionic Equation
- Cross Out Spectator Ions

**Don't you think we should try some examples? . . . . .**

Example 1: A solution of silver nitrate is combined with a solution of potassium chromate.

Example 2: A solution of lead(II)nitrate is added to a solution of potassium iodide (DEMO)

Example 3: Solutions of ammonium sulfate and potassium chloride are combined.