DESCRIPTIVE CHEMISTRY PRACTICE
(turning worded statements into chemical equations)

1. Mercury and bromine will react with each other to produce mercury (II) bromide.

2. Magnesium and Iodine will react with each other to produce magnesium iodide.

3. The reaction takes place when two reagents are combined, solid iron(III)oxide and aluminum metal. The products of this reaction are liquid (molten) iron and aluminum oxide.

4. The reaction between solid potassium chlorate and solid red phosphorus (P_4) takes place when you strike a match on a matchbox. The products of this reaction are the solids, tetraphosphorus decoxide and potassium chloride.

5. Solutions of sodium carbonate and silver nitrate react to form solid silver carbonate and a solution of sodium nitrate.

6. Solutions of sulfuric acid and lead(II)acetate react to form solid lead(II)sulfate and a solution of acetic acid.

Answers:

1. $Hg\ (l) + Br_2\ (l) \rightarrow HgBr_2\ (s)$

2. $Mg\ (s) + I_2\ (s) \rightarrow MgI_2\ (s)$

3. $Fe_2O_3\ (s) + 2\ Al\ (s) \rightarrow Al_2O_3\ (s) + 2\ Fe\ (l)$

4. $10\ KClO_3\ (s) + 3P_4\ (s) \rightarrow 3\ P_4O_{10}\ (s) + 10\ KCl\ (s)$

5. $Na_2CO_3\ (aq) + 2\ AgNO_3\ (aq) \rightarrow Ag_2CO_3\ (s) + 2\ NaNO_3\ (aq)$

6. $H_2SO_4\ (aq) + Pb(CH_3COO)_2\ (aq) \rightarrow PbSO_4\ (s) + 2\ CH_3COOH\ (aq)$